

# MARYLAND METRICS

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## TECHNICAL INFORMATION and DATA

Corrosive Agent	Concentration	L	Temperature	C	C + F	A2	A4	div.
Phosphoric acid (pure)  H <sub>3</sub> PO <sub>4</sub>	1 %		20°C	-	0	0	0	
	10 %		boiling	1	1	0	0	
			20°C	2	1	0	0	
	45 %		boiling	2	2	0	0	
			20°C	2	2	0	0	
	60 %		boiling	3	2	2	1	1+0*
			20°C	2	2	0	0	
	70 %		boiling	3	3	2	1	1+0*
		20°C	2	2	0	0		
80 %		boiling	3	3	2	2	1+0*	
		20°C	2	2	1	0		
		boiling	3	3	3	2	1+*	
concentrated			20°C	2	2	1	0	
			boiling	3	3	3	3	
Phosphorous pentoxide	P <sub>2</sub> O <sub>5</sub>	dry or humid	20°C	-	-	1	0	
Photographic developer	-	-	20°C	1	0	0	0	
Photographic fixer	-	x	20°C	3	3	1	0	
Picric acid	C <sub>6</sub> H <sub>2</sub> (NO <sub>2</sub> ) <sub>3</sub> OH	all concentrated	20°C	-	0	0	0	
Potassium acetate	CH <sub>3</sub> COO K	-	molten	-	-	0	0	
Potassium aluminium sulphate = alum KAL (SO <sub>4</sub> ) <sub>2</sub> , 12 H <sub>2</sub> O	10 %		20°C	1	0	0	0	
	hot saturated		boiling	2	2	1	0	
			20°C	2	2	0	0	
			boiling	3	3	3	2	1+
Potassium bichromate	K <sub>2</sub> Cr <sub>2</sub> O <sub>7</sub>	25 %	20°C	0	0	0	0	
			boiling	3	2	0	0	
Potassium bifluoride	KHF <sub>2</sub>	cold saturated	20°C	3	2	0	0	
Potassium bisulphate	K H SO <sub>4</sub>	2 %	90°C	-	-	3	2	0+
		5 %	20°C	-	-	1	0	
			90°C	-	-	3	2	0+^
		15 %	90°C	-	-	3	2	1+
Potassium bisulphite	Ca H <sub>2</sub> (SO <sub>3</sub> ) <sub>2</sub> = sulphite lye	-	20°C	2	2	0	0	
		20 at	boiling	3	3	2	0	
			200°C	3	3	3	0	
Potassium bitartrate	= tartar K H C <sub>4</sub> H <sub>4</sub> O <sub>6</sub>	hot saturated	cold boiling	- -	- -	0 2	0 1	0+ 0+
Potassium bromide	KBr	-	20°C	-	0	0	0	
Potassium chlorate	K Cl O <sub>3</sub>	hot saturated	boiling	-	0	0	0	
Potassium carbonate	K <sub>2</sub> CO <sub>3</sub> = potash	-	20°C	0	0	0	0	
			boiling	1	0	0	0	
Potassium chloride	Ca Cl <sub>2</sub> , 6 H <sub>2</sub> O	cold saturated	20°C	-	-	0	0	
			boiling	-	-	1	1	0°
Potassium chloride	K Cl	-	20°C	1	0	0	0	
		hot saturated	boiling	3	1	0	0	0°
Potassium chrome alum	= chrome alum KCr (SO <sub>4</sub> ) <sub>2</sub> , 12 H <sub>2</sub> O	-	20°C	2	2	0	0	0+
			boiling	3	3	3	3	1+
Potassium cyanate	K O C N	-	20°C	0	0	0	0	
Potassium cyanide	K C N	5 %	20°C	0	0	0	0	
Potassium cyanoferrate (III)	K <sub>3</sub> [Fe (CN) <sub>6</sub> ]	-	20°C	0	0	0	0	
		hot saturated	boiling	0	0	0	0	
Potassium cyanoferrate (II)	K <sub>4</sub> [Fe (CN) <sub>6</sub> ], 3 H <sub>2</sub> O	-	20°C and boiling	0	0	0	0	
Potassium hypochlorite	Ca (OCl) <sub>2</sub> , 4 H <sub>2</sub> O	cold saturated	to 40°C	-	-	2	1	0°
Potassium hydroxide	Ca (OH) <sub>2</sub> = slaked lime	-	20°C	0	0	0	0	
			boiling	-	-	0	0	
Potassium hydroxide = caustic potash  KOH	20 %		20°C	0	0	0	0	
			boiling	0	0	0	0	
			20°C	0	0	0	0	
			boiling	2	1	0	0	
		hot saturated fused mass	boiling	2	1	0	0	
			360°C	3	3	3	3	
Potassium hypochlorite	K Cl O	-	20°C	-	-	2	1	0°
			150°C	-	-	2	1	0°
Potassium iodide	KJ	-	20°C	2	1	0	0	
		x	boiling					
Potassium nitrate	= Kalisalpeter K NO <sub>3</sub>	25 %	20°C	0	0	0	0	
			boiling	-	0	0	0	
		50 %	20°C	0	0	0	0	
			boiling	-	0	0	0	
	molten mass	550°C	3	0	0	0		
Potassium oxalate	K <sub>2</sub> C <sub>2</sub> O <sub>4</sub> , H <sub>2</sub> O	all	20°C	0	0	0	0	
		concentrated	boiling	-	0	0	0	
Potassium permanganate	KMnO <sub>4</sub>	all	20°C	0	0	0	0	
		concentrated	boiling	3	1	0	0	
Potassium sulphate	K <sub>2</sub> SO <sub>4</sub>	-	20°C and boiling	-	0	0	0	